

**Bonds and Energy**

Use with textbook pages 123-124.

1. a) What type of compound are lithium fluoride and sodium oxide? Explain your answer.

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- b) What type of bond is formed between the atoms of this type of compound? Describe how it forms and how it affects stability.

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- c) Fill in the following table using Bohr diagrams to show how a lithium atom combines with a fluorine atom to form lithium fluoride. What is happening to the atoms and bonds in this reaction? Describe how energy is involved.

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lithium atom + fluorine atom	lithium fluoride
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- d) Fill in the following table using Bohr diagrams to show how sodium oxide is broken apart into a sodium atom and two oxygen atoms. What is happening to the atoms and bonds in this reaction? Describe how energy is involved.

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sodium oxide	sodium atoms + oxygen atom
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